


Chemical equilibrium worksheet

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Equilibrium Class 11 Chemistry Worksheet Pdf

Class 11 Chemistry students should refer to the following printable worksheet in Pdf for Equilibrium in standard 11. This test paper with questions and answers for Grade 11 Chemistry will be very useful for exams and help you to score good marks

Class 11 Chemistry Worksheet for Equilibrium

Equilibrium MCQ Questions with Answers

Class 11 Chemistry Question- In a vessel N₂, H₂ and NH₃ are at equilibrium. Some helium gas is introduced into the vessel so that total pressure increases while temperature and volume remain constant. According to Le Chatelier's principle, the dissociation of NH₃(a) increases(b) decreases(c) remains unchanged(d) equilibrium is disturbed

Answer-(c) Question-Which of the following molecules acts as a Lewis acid ?(a) (CH₃)₂O(b) (CH₃)₃P(c) (CH₃)₃N(d) (CH₃)₃B

Answer-(d) Question-Which of the following statements are correct regarding Arrhenius theory of acid and base?(a) This theory was applicable to only aqueous solutions(b) This theory was applicable to all solutions(c) This theory could not explain the basicity of substances like ammonia which do not possess a hydroxyl group(d) Both (a) and (c)

Answer-(d) Question-An acid/ base dissociation equilibrium is dynamic involving a transfer of proton in forward and reverse directions. Now, with passage of time in which direction equilibrium is favoured ?(a) in the direction of stronger base and stronger acid(b) in the direction of formation of stronger base and weaker acid(c) in the direction of formation of weaker base and weaker acid(d) in the direction of formation of weaker base and stronger acid

Answer-(c) Question-Would gaseous HCl be considered as an Arrhenius acid ?(a) Yes(b) No(c) Not known(d) Gaseous HCl does not exist

Answer-(b) Question-BF₃ is an acid according to(a) Arrhenius concept(b) Bronsted-Lowry concept(c) Lewis Concept(d) Both (b) and (c)

Answer-(c) Question- Which of the following is/are electrolytes?(i) Sugar solution (ii) Sodium chloride(iii) Acetic acid (iv) Starch solution(a) (i) and (iv)(b) (ii) and (iv)(c) (ii) and (iii)(d) (i) and (iii)

Answer-(c) Question-Effect of a catalyst on an equilibrium reaction.(i) A catalyst increases the rate of the chemical reaction by making available a new low energy pathway for the conversion of reactants to products.(ii) It increases the rate of forward and reverse reactions that pass through the same transition state and does not affect equilibrium.(iii) It lowers the activation energy for the forward and reverse reactions by exactly the same amount.

Which of the above statement(s) is/are correct ?(a) Only (i)(b) (i) and (ii)(c) (i), (ii) and (iii)(d) (ii) and (iii)

Answer-(c) Question-Conjugate acid of NH₂- NH is : (a) NH₄⁺(b) NH₃(c) NH₂(d) NH

Answer-(b) Question-Among boron trifluoride, stannic chloride and stannous chloride, Lewis acid is represented by(a) only stannic chloride(b) boron trifluoride and stannic chloride(c) boron trifluoride and stannous chloride(d) only boron trifluoride

Answer-(c) Question-Which one of the following molecular hydrides acts as a Lewis acid?(a) NH₃(b) H₂O(c) B₂H₆(d) CH₄

Answer-(c) Question-A base, as defined by Bronsted theory, is a substance which can(a) lose a pair of electrons(b) donate protons(c) gain a pair of electrons(d) accept protons

Answer-(d) Question- Which of the following can act as both Bronsted acid and Bronsted base?(a) Na₂CO₃(b) OH⁻(c) HCO₃⁻(d) NH₃

Answer-(c) Question-Which one of the following is the correct statement ?(a) HCO₃⁻ is the conjugate base of CO₃²⁻.(b) NH₂⁻ is the conjugate acid of NH₃.(c) H₂SO₄ is the conjugate acid of HSO₄⁻.(d) NH₃ is the conjugate base of NH₂⁻

Answer-(c) Question-The value of the ionic product of water(a) depends on volume of water(b) depends on temperature(c) changes by adding acid or alkali(d) always remains constant

Answer-(b) Question-A base when dissolved in water yields a solution with a hydroxyl ion concentration of 0.05 mol litre⁻¹. The solution is(a) basic(b) acidic(c) neutral(d) either (b) or (c)

Answer-(a) Question-The geometry of hydronium ion is(a) tetrahedral(b) linear(c) trigonal pyramidal(d) trigonal planer

Answer-(c) Question-pH scale was introduced by : (a) Arrhenius(b) Sorensen(c) Lewis(d) Lowry

Answer-(b) Question-Which of these is least likely to act as Lewis base?(a) F⁻(b) BF₃(c) PF₃(d) CO

Answer-(b) More Question

Instructions:1. All questions are compulsory.2. Please give the explanation for the answer where applicable.Q1 - (a) Write expression showing relationship between K_p and K_c for following reaction(b) Define conjugate acid and base with an example. Q2 - (i) Define the term 'pH of solution'. (ii) The hydrogen ion concentration of a solution is 10⁻⁴. Calculate the pH of solution. Q3 - At equilibrium, the concentrations of N₂=0.0032 M, O₂= 0.0043 M and NO =0.0026 M in a sealed vessel at 800K. What will be K_c for the reaction? Q4 - For the equilibrium, T he value of equilibrium constant, K_c is 4.30 x 10⁻⁶ at 1069 K. calculate the K_p for the reaction at this temperature? Q5 -Hydrolysis of sucrose gives, Sucrose + water Glucose + Fructose

Equilibrium constant, K_c for the reaction is 3x10¹¹ at 300 K. Calculate G at 300 K

Q6 -State Ostwald's dilution law.Q7 -The pK_a of acetic acid and pK_b of ammonium hydroxide are 4.82 and 4.72. Calculate the pH of ammonium acetate solution?Q8 - Calculate the solubility of AX in pure water. The solubility product of AX is 2.5 x10⁻²⁰. Please refer to the link below for CBSE Class 11 Chemistry Equilibrium Worksheet Set B

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